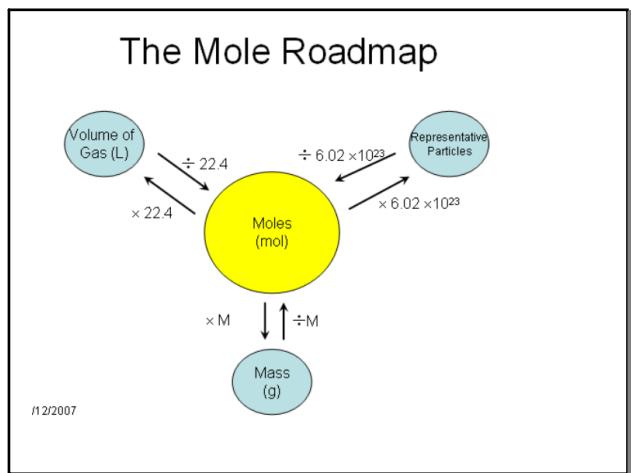
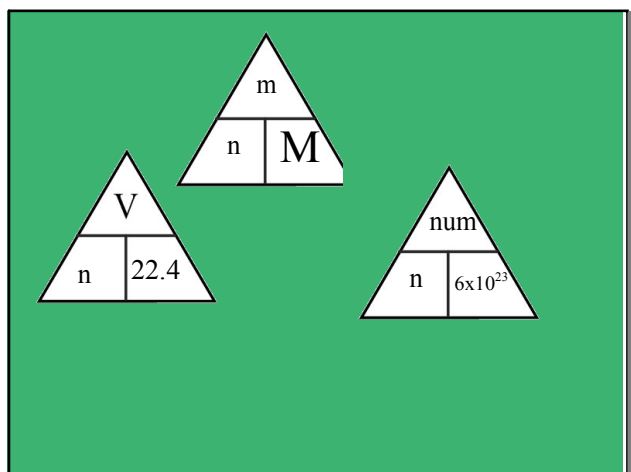


Nov 12-1:59 PM



Nov 12-2:02 PM



Nov 12-2:03 PM

You need to become familiar with the following types of calculations

- 1. mole > mole
- 2. mole > gas volumes (STP + NTP)
- 3 mole > mass
- 4 mass > mole
- 5 mass > mass



easiest
hardest, most steps

Nov 12-2:05 PM

Some hand written examples

- 1. mole > mole

Nov 12-2:12 PM

Some hand written examples

- 1. mole > mole

Nov 12-2:12 PM

Some hand written examples

2. mole > gas volumes (STP + NTP)

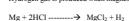
Hydrogen gas is produced when magnesium is added to dilute hydrochloric acid.



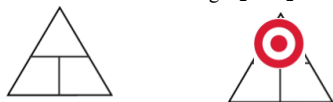
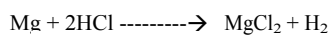
What volume of gas is produced when 0.6 moles of the acid is used completely?

Nov 12-2:12 PM

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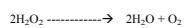


Nov 12-2:12 PM

Some hand written examples

2. mole > gas volumes (STP + NTP)

The lab prep for oxygen requires hydrogen peroxide and the catalyst manganese dioxide



25mL. of oxygen gas is produced during an experiment. How many moles of hydrogen peroxide has decomposed?




Nov 12-2:12 PM

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$$2\text{H}_2\text{O}_2 \text{ -----} \rightarrow 2\text{H}_2\text{O} + \text{O}_2$$

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Nov 12-2:12 PM

Some hand written examples






2. mole > gas volumes (STP + NTP)

Carbon dioxide is made by adding hydrochloric acid to calcium carbonate. A teacher requires 150mL at NTP. How g of calcium carbonate should he use?

$$\text{CaCO}_3 + 2\text{HCl} \text{ -----} \rightarrow \text{CaCl}_2 + \text{H}_2\text{O} + \text{CO}_2$$

Nov 12-2:12 PM

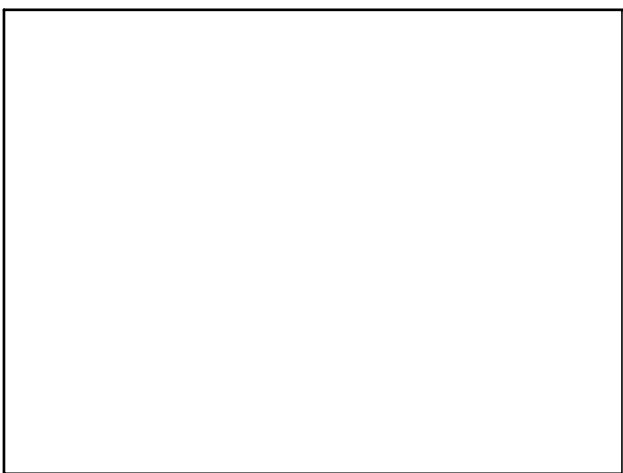
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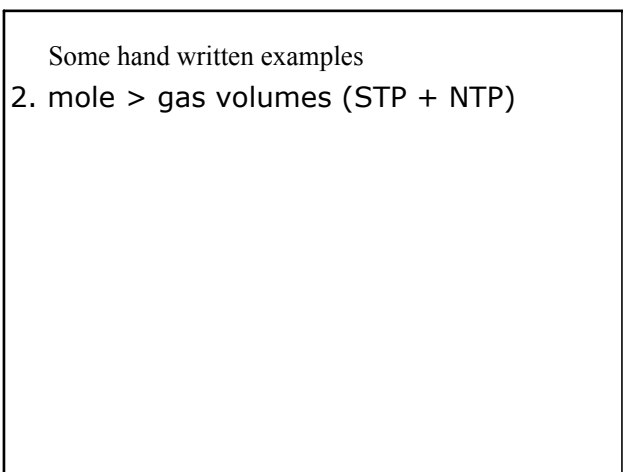
Nov 12-2:12 PM



Nov 16-7:37 AM



Nov 12-2:12 PM



Nov 12-2:12 PM

Some hand written examples

3 mole > mass

Nov 12-2:12 PM

Some hand written examples

3 mole > mass

Nov 12-2:12 PM
